

Conservation 1404 - Review of Chemistry of Pollution
Pages 15-16, 35 – 36, 40, and fig. 2.5
“The Least You Should Know”

Vocabulary:

Pollution
Point sources
Non-point sources
Pollution prevention = Input pollution control
Pollution cleanup = Output pollution control
Ion
pH, neutral, acidic, basic
Molecule
Chemical formula
Concentration in ppm = parts per million
Persistence (of a pollutant)

Concepts:

Pollutants can enter the environment naturally, and know some common mechanisms.
Pollutants contaminate the areas where they are produced but wind and flowing water can carry them to other areas.
Nonpoint sources are difficult to identify and give examples.
Eroded sediment can be a source of pollution
Unwanted effects of pollution:
 → disrupt or degrade life support systems
 → directly damage health and property
 → create nuisances such as smells, tastes, sights.
Basic structure of an atom
Basic structure of a monatomic ion
How to read the charges on an ion, for example Ca^{2+}
How to read the charges on polyatomic ions, e.g. SO_4^{2-}
Be familiar with fig. 2.5 in general terms & stomach acid, lemons, rainwater, blood, phosphate detergents, ammonia.

How to read chemical formulas
Organic vs inorganic compounds
Persistence of a pollutant and examples of each:
 Degradable pollutants
 Biodegradable pollutants
 Slowly degradable pollutants
 Nondegradable pollutants

